

Initial development of a linear regression model to determine the copper (ii) ion content via a photometric method

Nguyen Thi Anh Hong*, Nguyen Thi Thuy Dung, Nguyen Pham Thien Thanh



Use your smartphone to scan this QR code and download this article

ABSTRACT

Introduction: This study, which was first conducted in Vietnam, aimed to develop a multivariable and simple-variable linear regression model from the direct measurement of the UV–Vis absorption of copper(II) ions in aqueous solution without using other reagents (chelating agents and solvents), which reduces environmental pollution and analysis fees. **Methods:** Simple-variable and multivariable linear regression models were developed from UV–Vis spectral data of copper(II) ion solutions with concentrations ranging from 0.2 to 50 ppm. **Results:** Four multivariable regression models were developed and modified, and the optimal simple variable regression model was selected. This study analyzed the suitability of single and multivariable models for the analysis of copper(II) ions in aqueous solution at low concentrations. **Conclusion:** This study successfully built and adjusted linear regression models for predicting the copper(II) ion content in aqueous solution via a photometric method. The multivariable model with odd variables (model No. 2') and the simple-variable model at a wavelength of 221 were optimized for use in the prediction of the concentration at an acceptable level of 0.5 ppm. These results were verified by the graph of the correlation between the true concentration and the predicted concentration in both selected models. In particular, the multivariate model yields significantly more accurate prediction results than does the simple-variable model.

Key words: Copper(II) ion, linear regression model, photometric method

Department of Chemistry, College of Natural Sciences, Can Tho University, Can Tho, Viet Nam

Correspondence

Nguyen Thi Anh Hong, Department of Chemistry, College of Natural Sciences, Can Tho University, Can Tho, Viet Nam
Email: ntahong@ctu.edu.vn

History

- Received:
- Accepted:
- Published Online:

DOI :



Copyright

© VNUHCM Press. This is an open-access article distributed under the terms of the Creative Commons Attribution 4.0 International license.



1 INTRODUCTION

The concentration of heavy metal ions is usually determined by methods such as complexometric titration, voltammetric methods, and photometric methods (UV–Vis), among others. Atomic absorption spectrometry (AAS)¹ or electrothermal atomic absorption spectrometry (ETAAS) are among the most commonly used techniques for trace element analyses of minerals.² In addition, there are more sophisticated methods, such as inductively coupled plasma optical emission spectroscopy (ICP–MS), which uses plasma to analyze trace metal ions in beverages,³ and high-resolution inductively coupled plasma–mass spectrometry (HR–ICP–MS), which has an electrical and magnetic region for ion separation and concentration in industrial wastewater analysis.⁴ These analytical methods provide high selectivity, high sensitivity and low detection limits, but the equipment is complicated and quite expensive. For example, Huang and Shih (1993) directly detected copper in seawater samples using a graphite furnace atomic absorption spectrometer (GFAAS) with high accuracy and precision to detect Cu(II), where the detection limit of Cu(II) was in the range of 0.3–0.4 $\mu\text{g. L}^{-1}$

when injected with 20 μl of seawater, which further decreased to 0.07 $\mu\text{g. L}^{-1}$ with multiple injections.⁵ An optimized single-particle ICP–MS technique (spICP–MS) was used by Venkatesan et al. (2018) to analyze Pb, Fe, Sn, Cu, and Ag in tap water samples. spICP–MS is a time-resolved analysis in which particles are detected as collisions above the elemental background signal. This instrument detected Cu(II) in 25 water samples in the temperature range of 15–136 g.L^{-1} .⁶

In Vietnam, the fluorescent chemosensor in the UV–Vis machine has been researched by Duong Tuan Quang and his associates since 2007 with a number of publications, such as a chemical sensor based on calix[4]arene to detect the ions Fe^{3+} , F^{-} , Cs^{+} , Cu^{2+} ; or dimethylaminocinnamaldehydeaminothiourea to detect Ag^{+} , Cu^{2+} , Hg^{2+} ; and a chemical sensor containing a 1,2,3-triazole ring that detects Al^{3+} or chemical sensors that detect Hg^{2+} synthesized from rhodamine derivatives or fluorescent reagents.^{7–10}

This photometric method requires no user training, simple equipment, and easy sample handling. Currently, in Vietnam and around the world, copper(II) ions are analyzed by this method, and researchers use reagents in combination with copper(II) ions to form

Cite this article : Hong N T A, Dung N T T, Thanh N P T. Initial development of a linear regression model to determine the copper (ii) ion content via a photometric method. *Sci. Tech. Dev. J.* 2024; ():1-11.

complexes whose color is detected in the Vis region. A study by Sharma et al. (2010) showed that copper can be detected at the maximum absorption wavelength (336 nm) through the use of a novel UV spectroscopy method (Shimadzu UV-Visible 160 A spectrometer) based on the formation of complexes of Cu(II) ions with cefixime immediately in a 1,4-dioxan-distilled water medium at room temperature. In this study, the proposed method was able to analyze Cu(II) in natural water samples with a detection limit of $3.19 \times 10^{-2} \mu\text{g/mL}$.¹¹ The reagent 1-(2-pyridylazo)-2-naphthol (PAN) was used to analyze copper in sugarcane spirit. Complexation at pH 4.50 for 5 min at 20 °C requires a malonic acid coating to reduce the influence of iron(III) and nickel(II) ions. Linearity was obtained with a copper(II) concentration of 8.00 mg/L, and the limits of detection and quantification were 0.02 mg/L and 0.13 mg/L, respectively.¹² In 2012, Omar et al. adopted near-infrared spectroscopy analysis in the commonly used 700 to 1100 nm range to reliably determine the dissolved solids content in fruit. The aim was to optically profile the sugar-water solution and determine the peak wavelength in the quantification of the sugar concentration.¹³ This method was developed for the analysis of metal ions (copper and lead) in aqueous solution in 2014 by Tan's research group at the University of Sains, Malaysia. This research group has produced multivariable and simple-variable linear regression models for the analysis of metal ions at low concentrations from 0.2 to 10 ppm by photometric methods without using any reagents.¹⁴ This is the initial study of a multivariable and simple-variable linear regression model for the direct measurement of the UV-Vis absorption of copper(II) ions in aqueous solution without the use of other reagents (chelating agents and solvents), which reduces environmental pollution and analysis fees. This study was first conducted in Vietnam with the desire to contribute to the expansion of analytical methods that do not use chemical reagents.

MATERIALS AND METHODS

General information

All chemicals used in this study were of analytical grade. UV-Vis spectra were measured using a UV-Vis instrument (Jasco V-730). The data were analyzed by Microsoft Excel.

General method for preparation of samples

For the stock standard solution [Cu] at 1000 ppm, 2.683 mg of $\text{CuCl}_2 \cdot 2\text{H}_2\text{O}$ was accurately weighed into a 100

mL beaker to dissolve enough distilled water, after which the solution in the beaker was transferred to a 1 L volumetric flask, and distilled water was added. The flask was closed tightly and shaken by inverting several times until the solution was homogeneous. *The intermediate standard solution:* From the 1000 ppm stock standard, a series of intermediate standards with concentrations of 5 ppm, 10 ppm, 20 ppm, 30 ppm, 40 ppm and 50 ppm were prepared. *Low-concentration standard solution:* From the 5 ppm intermediate standard solution, a series of intermediate standards with concentrations of 0.2 ppm, 0.5 ppm, 1 ppm and 2 ppm were prepared.

Experiment

Intermediate standard solutions of 5-50 ppm concentration were used to measure the absorbance in triplicate. The results of the spectrum were used to develop the regression model. Standard solutions with low concentrations of 0.2-2 ppm were measured for absorbance in triplicate. The results of the spectra were subjected to the optimized regression model to calculate the amount of copper(II) ions.

Data processing

Determine the appropriate wavelength range: To select the appropriate wavelength for the linear regression model (LRM) structure, the noisy and near-baseline regions need to be removed.

Multivariable LRM: After wavelength selection, the data exported from the spectra were analyzed via multivariate LRM with the proposed models. Next, we identify the independent variables that have a weak correlation with the dependent variable and remove them. The multivariable LMR run was repeated with the remaining variables, and the linear regression equation (LRE) was determined. There are four proposed multivariate LRMs:

- Model No. 1: The selected variables have values of 1 wavelength apart;
- Model No. 2: The selected variables have even wavelengths;
- Model No. 2': The selected variables have odd wavelengths;
- Model No. 3: The selected variables have 5 different wavelengths.

Design of the simple-variable LRM: After optimization, the simple-variable LMR for each variable corresponding to different wavelengths is analyzed for the variables selected in the multivariable model.

By applying these models to solutions of low-concentration standard solutions, the data exported from the spectra were analyzed by multivariate and simple variable LRM, which were optimized.

Conditions for satisfying the optimal model

The optimal model is the model with no more than 5 independent variables; $0.99 \leq R^2 \leq 1$ and adjusted R^2 between 0.5 – 1; small standard error - error (10^{-3}); absolute deviation - bias (%) < 15% (according to many organizations in the US, Canada, Europe – ISO 3534-1)

RESULTS

The appropriate wavelength range

The experiment was conducted using a two-channel spectrometer with wavelengths ranging from 200 nm to 1100 nm. However, the results show that channel 1 (wavelengths of 650 nm to 1100 nm) is not Visible at low concentrations of Cu^{2+} . Moreover, measurements through channel 0 (200 nm to 650 nm) produced a significant coefficient of determination, R^2 , between the absorbance and copper ion concentration. Spectroscopic results in the 200-230 nm working region show that the data at wavelengths below 217 nm are noisy (Figure 1). Therefore, the extreme negative peak near 217 nm was neglected in this study.

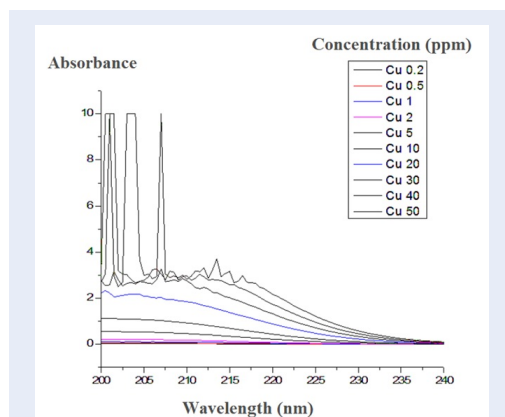


Figure 1: The appropriate wavelength range of copper(II) ion solutions with concentrations ranging from 0.2 to 50 ppm

174

The multivariable LRM

Model No. 1

In the selected working area from 217-230 nm, the regression coefficients of wavelengths 217, 219, 220,

223, 226, 227, 229 and 230 and the intercept show a weak correlation with the regression equation and should be rejected.

The result of the model No. 1 modification is linear regression equation-1 $C = 9.5 \times D_{218} - 29.6 \times D_{221} - 8.8 \times D_{222} + 105.3 \times D_{224} - 15.1 \times D_{225} - 36.5 \times D_{228}$ (LRE-1), with a mean error and absolute deviation of 0.711 and 0.77, respectively (Table 1). LRE-1 had a lower adjusted coefficient of determination ($R^2_{adj.} = 0.89998$) than the original regression equation (0.99987) but is still quite good for the linear regression method.

Models No. 2 and No. 2'

For Model No. 2, the intercept and regression coefficients at 218, 220 and 222 nm show a weak correlation with the model, so they are ignored. After recalibration, this model gives the equation $C = 95.2 \times D_{224} - 38.9 \times D_{226} - 96.6 \times D_{228} + 60.7 \times D_{230}$ (LRE-2). However, at odd wavelengths (model No. 2'), the model is corrected after removing the variable with a weak correlation at wavelengths 217 and 227 nm, and the equation $C = 11.5 \times D_{219} - 23.9 \times D_{221} + 78.6 \times D_{223} - 13.1 \times D_{225} - 56.4 \times D_{229}$ (LRE-3) is obtained. The adjusted coefficients of determination for both models (No. 2 and No. 2') are 0.92850 and 0.92302, respectively.

Table 2 shows that both models are suitable for predicting analyte content; however, model No. 2' is superior when it has a relatively small error, approximately 2.144×10^{-3} .

Model No.3

The number of independent variables in this model is the lowest when the independent variables have a long jump (5 nm). After adjusting to remove the weak correlation to the dependent variable of the constant, this model has 3 independent variables with $C = 11.9 \times D_{220} + 33.2 \times D_{225} - 29.2 \times D_{230}$ (LRE-4) and $R^2_{adj.} = 0.93318$.

This model has the advantage of the number of independent variables, but the results of the analysis of the parameters (Table 3) show that the error of the model compared to the real value is quite large, up to 30.7×10^{-3} ; therefore, the forecasting results are not as good as those of the above models.

The simple-variable LRM

The variables selected in Model adjusted No. 2 have variables at 219, 221, 223, 225 and 229 nm. At each of these wavelengths, simple-variable regression analysis is performed and modified when the intercept is not significant for the model.

Table 1: The results of the error analysis and absolute bias of adjusted model No.1

True C (ppm)	Pred. C (ppm)	Error (10 ⁻³)	Bias (%)
5	5.09	0.478	1.85
5	5.33	6.056	6.60
5	5.05	0.148	1.03
10	10.01	0.005	0.10
10	9.89	0.708	1.13
10	10.00	0.001	0.05
20	20.05	0.152	0.26
20	19.90	0.607	0.52
20	20.23	2.866	1.14
30	29.85	1.268	0.50
30	29.96	0.079	0.13
30	30.02	0.018	0.06
40	40.01	0.009	0.03
40	39.96	0.083	0.10
40	40.02	0.023	0.05
50	50.04	0.087	0.08
50	49.94	0.186	0.12
50	50.02	0.028	0.05
Average		0.711	0.77

The absolute deviation of concentrations in simple-variable regression analysis at different wavelengths shows that lower concentrations (5 and 10 ppm) have much larger deviations at higher concentrations (Figure 2). In Table 4, the results of the error analysis of the variables 219, 225 and 229 show much larger values (14.1, 14.0 and 17.2, respectively) than those at wavelengths 221 and 223 (6.4 and 8.7, respectively). The absolute and absolute deviations are similar but significantly smaller at these two wavelengths (1.18% and 1.31%, respectively, at 221 and 223). Therefore, the optimal model is selected at these two wavelengths because the results of the error and absolute deviation analysis have lower values at wavelength 221. Therefore, the optimal simple-variable LRE is developed in this model, $C = 24.923 \times D221$ (LRE-5), corresponding to an R^2 of 0.99987.

In conclusion, model-adjusted No. 2' was chosen as the optimized model for multivariable linear regression analysis, with $C = 11.5 \times D219 - 23.9 \times D221 + 78.6 \times D223 - 13.1 \times D225 - 56.4 \times D229$ (LRE-3) and simple-variable LRM $C = 24.923 \times D221$ (LRE-5).

Applying the optimal models to solutions with low-concentration standard solutions

The absorbance results of the low-concentration standard solutions with a concentration of 0.2-2 ppm were applied to the optimal models, and the results are summarized in Table 5.

To determine whether the simple-variable LRM is optimized at a wavelength of 221, this method was applied at each wavelength for concentrations ranging from 0.2-2 ppm, and the results are shown in Table 6. When analyzing data for the absolute deviation of wavelengths at low concentrations, the results are similar to those of the multivariable regression model (2' model), and simple variable models can only be properly applied to concentrations above 0.2 ppm (Table 6). From the graph showing the absolute deviation at low concentrations with different wavelengths, the density of the model at wavelength 221 is very high (Figure 3), which proves that simple variable information selection is effective at 221 nm. The analytical error and absolute deviation values in Table 6 also support this choice.

Table 2: The results of the error analysis and absolute bias of adjusted models No. 2 and No. 2'

True C (ppm)	Model No.2		Model No.2'	
	Error (10^{-3})	Bias (%)	Error (10^{-3})	Bias (%)
5	4.704	5.8	0.148	1.0
5	1.718	3.5	5.021	6.0
5	0.157	1.1	3.668	5.1
10	0.357	0.8	0.655	1.1
10	1.782	1.8	0.613	1.1
10	0.924	1.3	6.227	3.3
20	8.081	1.9	0.001	0.0
20	0.048	0.1	0.518	0.5
20	4.347	1.4	6.693	1.7
30	8.569	1.3	0.062	0.1
30	0.046	0.1	4.825	1.0
30	0.312	0.2	0.710	0.4
40	1.649	0.4	0.535	0.2
40	0.024	0.1	7.438	0.9
40	0.339	0.2	1.115	0.4
50	2.026	0.4	0.210	0.1
50	1.358	0.3	0.118	0.1
50	11.667	0.9	0.026	0.0
Average	2.673	1.2	2.144	1.3

When using the Optimum LRMs to predict copper(II) ions at low concentrations, the analytical results show that the acceptable concentration for this model is no less than 0.5 ppm.

DISCUSSIONS

All of the modified models have adjusted R^2 and R^2 values close to 1 (0.999999), so they are suitable for the requirements set for model selection. Model No. 1 has a large number of independent variables, making it difficult to predict. Although Model No. 3 has the fewest variables, it has a much larger forecast error than the other models. For models No. 2 and No. 2', when comparing the errors and bias, the 2' model is more suitable for choosing the optimal model. The result of selecting the multivariable regression model was Model No. 2' with 5 independent variables: 219, 221, 223, 225 and 229. In addition, the simple-variable regression model was selected with the same wavelength of 221 as Tan's model but dif-

ferent regression coefficients (24.923 and 79.311, respectively).¹⁴ When two optimization models were applied to predict copper(II) ions at low concentrations, the results were similar to those of Tan's model, which could predict concentrations of approximately 0.5 ppm (Table 6).

Verification of the simple-variable LRM at 221 with LRE-5 was performed by calculating analytic concentrations from 0.5 to 50 ppm and then graphing the correlation between the predicted and true concentrations. The analysis results are shown in the graph (Figure 4) and equation $y = 0.9999x$, which proves that the model used to predict the results is similar (99.99%) to the real value. In addition, this graph clearly shows that the multivariate model (model No. 2') is significantly better suited for forecasting than the simple-variable model with the correlation function $y = x$.

CONCLUSION

This study successfully developed simple-variable and multivariable linear regression models for copper(II)

Table 3: The results of the error analysis and absolute bias of adjusted model No. 3

True C (ppm)	Pred. C (ppm)	Error (10 ⁻³)	Bias %
5	5.018	0.018	0.36
5	4.974	0.038	0.52
5	5.282	4.421	5.64
10	10.036	0.072	0.36
10	9.770	2.928	2.30
10	10.360	7.193	3.60
20	19.991	0.004	0.04
20	19.720	4.349	1.40
20	20.380	8.041	1.90
30	29.956	0.107	0.15
30	29.657	6.539	1.14
30	30.549	16.756	1.83
40	40.424	9.989	1.06
40	39.256	30.711	1.86
40	40.041	0.094	0.10
50	50.100	0.552	0.20
50	49.408	19.473	1.18
50	50.476	12.579	0.95
Average		6.881	1.37

ion concentrations in aqueous solutions ranging from 0.2-50 ppm without using any other reagents or solvents in the wavelength range 217-230 nm. The results show that the multivariable model with odd variables (model No. 2') and the simple-variable model at a wavelength of 221 were optimized for use in predicting the concentration at an acceptable level of 0.5 ppm. These results were verified by the graph of the correlation between the true concentration and the predicted concentration in both selected models. In particular, the multivariate model yields significantly more accurate prediction results than does the simple-variable model.

The results of this study show that the application of multivariate and simple-variable regression models can almost accurately predict low copper(II) ion concentrations (0.5-50 ppm). However, the suitability of the models for analyzing complex samples and the factors affecting the analysis results, such as pH and metal ions, has not yet been investigated. For further research, this technique can be simplified to a more portable device at a lower cost using modern equipment.

COMPETING INTERESTS

The authors declare that they have no competing interests.

AUTHOR CONTRIBUTIONS

Nguyen Thi Anh Hong conceived the idea and designed the works. Nguyen Thi Thuy Dung and Nguyen Pham Thien Thanh performed experiments. All authors analyzed data, read and final approval manuscript for publication.

ACKNOWLEDGMENT

I would like to express my deepest gratitude to the Department of Chemistry, Faculty of Natural Sciences, Can Tho University for creating favorable conditions for us to complete this research.

ABBREVIATIONS

- L is a metric unit of volume (Liter)
- LRE The linear regression equation
- LRM The linear regression model
- ppm Parts per million corresponds to mg/L

Table 4: The results of the error analysis and absolute bias for model No. 2' and the simple-variable LRM at different wavelengths

True C (ppm)	D219			D221			D223			D225			D229			Model No.2'		
	Error (10 ⁻³)	Bias (%)	Error (10 ⁻³)	Error (10 ⁻³)	Bias (%)	Error (10 ⁻³)	Error (10 ⁻³)	Bias (%)	Error (10 ⁻³)	Error (10 ⁻³)	Bias (%)	Error (10 ⁻³)	Error (10 ⁻³)	Bias (%)	Error (10 ⁻³)	Error (10 ⁻³)	Bias (%)	Error (10 ⁻³)
5	1.706	3.50	0.026	0.089	0.43	0.089	0.089	0.80	0.001	0.10	4.953	0.148	5.97	1.03	0.655	6.01	5.14	0.001
5	0.882	2.52	0.097	0.016	0.84	0.016	0.016	0.34	0.367	1.63	3.940	0.655	5.33	6.01	0.001	5.14	1.09	0.062
5	4.374	5.61	1.112	2.726	2.83	2.726	2.726	4.43	3.371	4.93	5.269	0.001	6.16	5.14	0.001	5.14	1.09	0.062
10	1.804	1.80	0.045	0.260	0.28	0.260	0.260	0.68	1.401	1.59	8.509	0.062	3.91	1.09	0.062	1.09	1.05	0.535
10	1.161	1.45	1.134	3.202	1.43	3.202	3.202	2.40	7.255	3.61	12.703	0.535	4.78	1.05	0.535	1.05	1.05	0.535
10	14.232	5.06	3.671	2.614	2.57	2.614	2.614	2.17	5.419	3.12	1.898	0.210	1.85	3.35	0.210	3.35	3.35	0.210
20	2.263	1.01	2.556	3.439	1.07	3.439	3.439	1.24	5.944	1.64	3.736	5.021	1.30	0.02	5.021	0.02	0.02	5.021
20	2.461	1.05	3.336	6.781	1.23	6.781	6.781	1.75	11.229	2.25	20.532	0.613	3.04	0.48	0.613	0.48	0.48	0.613
20	30.816	3.72	5.862	3.086	1.62	3.086	3.086	1.18	8.374	1.94	4.586	0.518	1.44	1.74	0.518	1.74	1.74	0.518
30	10.272	1.43	1.059	0.011	0.46	0.011	0.011	0.05	0.051	0.10	13.919	4.825	1.67	0.11	4.825	0.11	0.11	4.825
30	1.040	0.46	23.255	28.950	2.16	28.950	28.950	2.41	29.845	2.44	48.832	7.438	3.13	0.98	7.438	0.98	0.98	7.438
30	39.869	2.82	9.362	3.094	1.37	3.094	3.094	0.79	10.376	1.44	16.294	0.118	1.81	0.38	0.118	0.38	0.38	0.118
40	17.832	1.42	1.733	0.018	0.44	0.018	0.018	0.05	0.377	0.21	2.884	3.668	0.57	0.25	3.668	0.25	0.25	3.668
40	5.660	0.80	27.390	27.476	1.76	27.476	27.476	1.76	19.964	1.50	45.316	6.227	2.26	0.91	6.227	0.91	0.91	6.227
40	8.775	0.99	0.263	2.795	0.17	2.795	2.795	0.56	7.114	0.89	2.089	6.693	0.48	0.35	6.693	0.35	0.35	6.693
50	2.948	0.46	10.381	0.890	0.86	0.890	0.890	0.25	35.009	1.59	15.270	0.710	1.05	0.12	0.710	0.12	0.12	0.710
50	2.113	0.39	5.730	3.206	0.64	3.206	3.206	0.48	2.545	0.43	7.391	1.115	0.73	0.09	1.115	0.09	0.09	1.115
50	105.500	2.76	17.492	67.559	1.12	67.559	67.559	2.21	102.869	2.72	90.813	0.026	2.56	0.04	0.026	0.04	0.04	0.026
Average	14.095	2.07	6.361	8.678	1.18	8.678	8.678	1.31	13.973	1.78	17.163	2.144	2.67	1.29	2.144	1.29	1.29	2.144

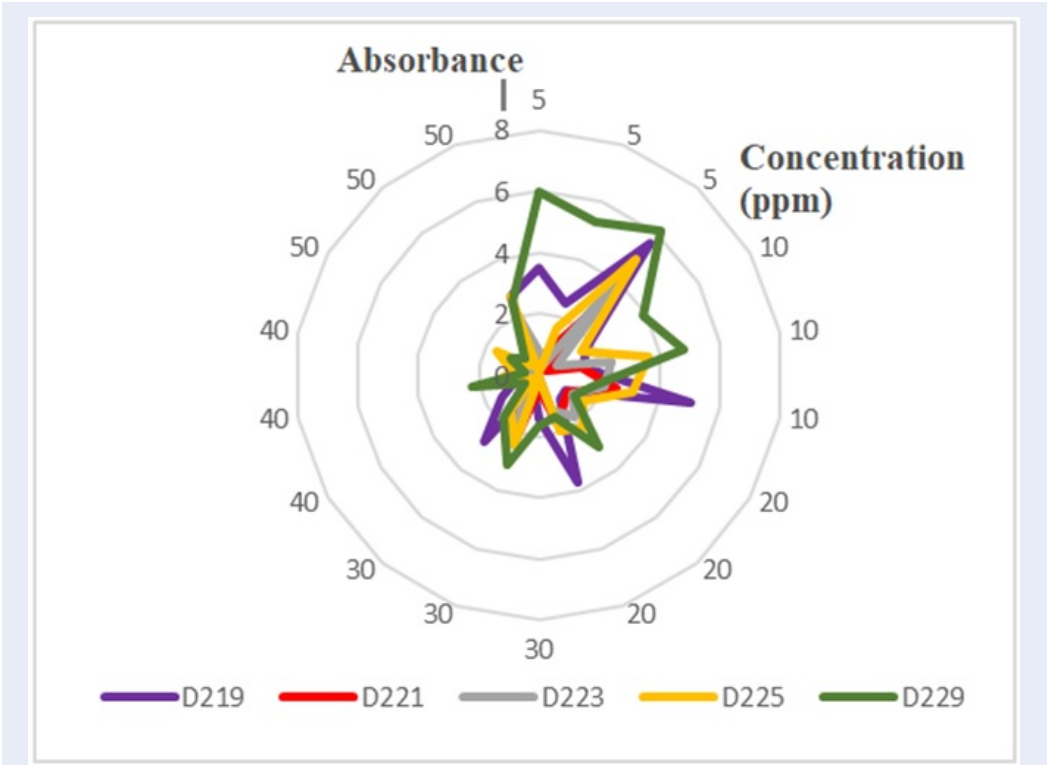


Figure 2: The absolute deviation at different wavelengths of copper(II) ion solutions with concentrations ranging from 5 to 50 ppm (the vertical numbers are Abs, and the numbers around the circle are the concentrations of Cu²⁺).

Table 5: The results of the error analysis and absolute bias of adjusted No. 2' at low concentrations

True C (ppm)	Pred. C (ppm)	Error (10-3)	Bias (%)
0.2	0.267	0.252	33.7
	0.135	0.232	32.3
	0.263	0.220	31.4
0.5	0.482	0.019	3.7
	0.470	0.049	5.9
	0.470	0.050	6.0
1.0	0.903	0.527	9.7
	0.948	0.152	5.2
	0.903	0.523	9.7
2.0	2.190	2.005	9.5
	1.973	0.041	1.4
	2.134	0.991	6.7

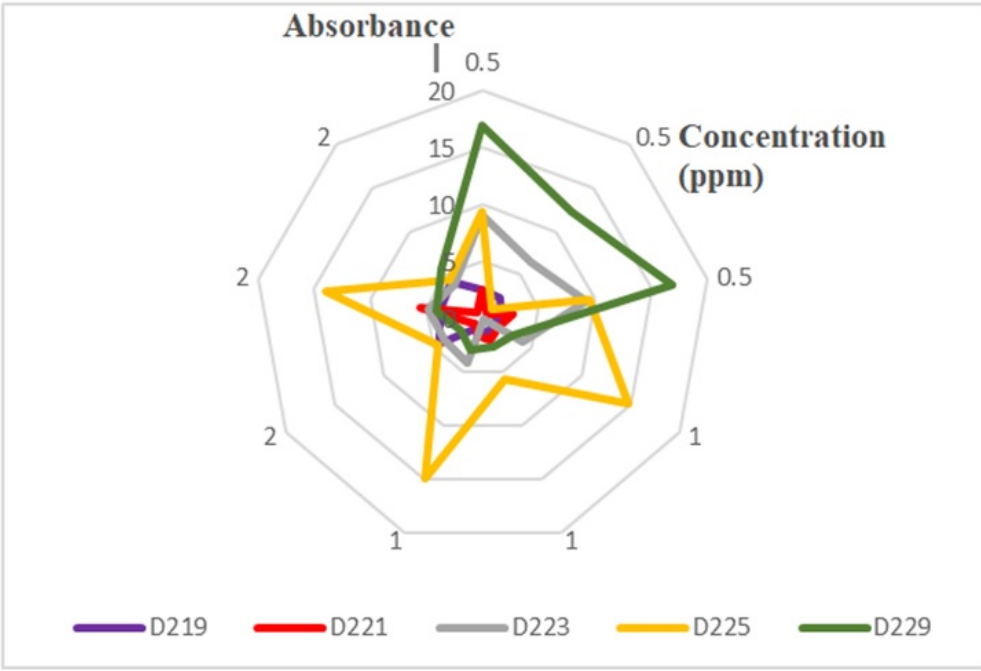


Figure 3: The absolute deviation at different wavelengths of copper(II) ion solutions with concentrations ranging from 0.5 to 2 ppm (the vertical numbers are Abs, and the numbers around the outer ring are the concentrations of Cu²⁺).

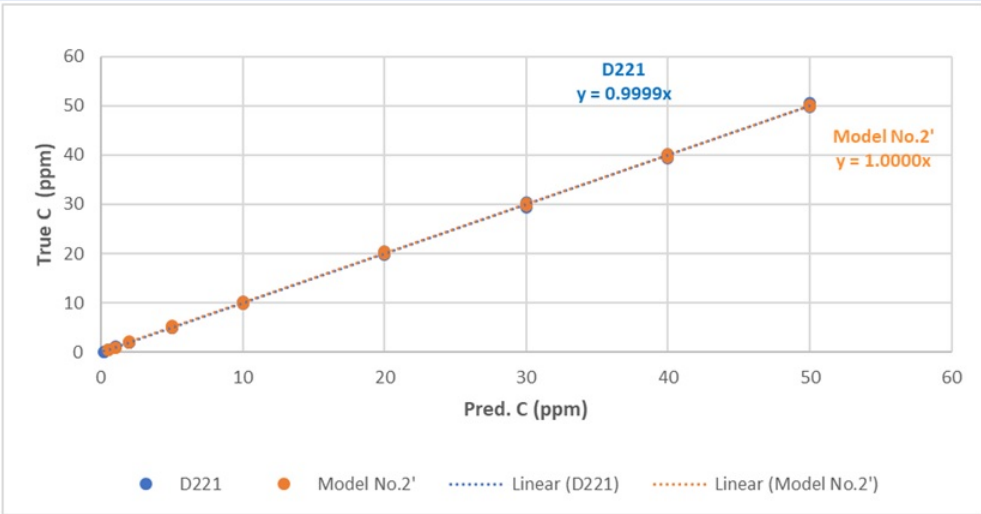


Figure 4: Correlation between the true and predicted concentrations of the simple-variable LRM at 221 nm and model No. 2'

Table 6: The results of the analysis of the parameters at low concentrations

True C (ppm)		0.2	0.5	1.0	2.0
D219	LRE	C = 20.589 x Abs ($R^2 = 0.99972$)			
	Pred. C (ppm)	0.045	0.505	1.000	2.032
	Error (10^{-3})	1.339	0.007	0.007	0.387
	Bias %	77.5	2.1	1.1	4.2
D221	LRE	C = 24.923 x Abs ($R^2 = 0.99987$)			
	Pred. C (ppm)	0.132	0.492	1.017	1.953
	Error (10^{-3})	0.259	0.006	0.017	0.227
	Bias %	34.1	1.9	1.7	2.3
D223	LRE	C = 31.496 x Abs ($R^2 = 0.99983$)			
	Pred. C (ppm)	0.147	0.459	0.972	2.020
	Error (10^{-3})	0.187	0.094	0.063	0.376
	Bias %	26.3	8.1	2.8	4.1
D225	LRE	C = 40.830 x Abs ($R^2 = 0.99972$)			
	Pred. C (ppm)	0.062	0.470	0.924	1.973
	Error (10^{-3})	1.058	0.069	0.810	1.649
	Bias %	69.0	6.1	11.4	7.6
D229	LRE	C = 72.973 x Abs ($R^2 = 0.99966$)			
	Pred. C (ppm)	0.032	0.428	1.026	2.031
	Error (10^{-3})	1.567	0.289	0.039	0.399
	Bias %	84.0	14.3	2.6	4.0

R2adj. The adjusted coefficient of determination

UV-Vis Ultraviolet visible spectrophotometers

REFERENCES

- Zhang Y, Qiao R, Sheng C, Zhao H. Chapter 4 - Technologies for detection of HRP in wastewater. High-Risk Pollutants in Wastewater. 2020 Jan 1:79-100; Available from: <https://www.sciencedirect.com/science/article/pii/B9780128164488000046>.
- Holcombe JA, Borges DLG. Graphite Furnace Atomic Absorption Spectrometry. Encyclopedia of Analytical Chemistry. 2010 March 15; Available from: <https://doi.org/10.1002/9780470027318.a5108.pub2>.
- Chen C-Y, Aggarwal SK, Chung C-H, You C-F. 7 - Advanced Mass Spectrometry for Beverage Safety and Forensic. Safety Issues in Beverage Production. 2020 Jan 1:223-69; Available from: <https://www.sciencedirect.com/science/article/pii/B9780128166796000073>.
- Einschlag FSG, Carlos L. Waste water: treatment technologies and recent analytical developments. 2013 Jan 16;.
- Huang S-D, Shih K-Y. Direct determination of copper in sea-water with a graphite furnace atomic absorption spectrometer. Spectrochimica Acta Part B: Atomic Spectroscopy. 1993 Oct 1; 48(12):1451-60; Available from: <https://www.sciencedirect.com/science/article/pii/S058485479380133F>.
- Venkatesan AK, Rodríguez BT, Marcotte AR, Bi X, Schoepf J, Ranville JF, et al. Using single-particle ICP-MS for monitoring metal-containing particles in tap water. Environmental Science: Water Research & Technology. 2018; 4(12):1923-32; Available from: <http://dx.doi.org/10.1039/C8EW00478A>.
- Kim JS, Quang DT. Calixarene-Derived Fluorescent Probes. Chemical Reviews. 2007 Sep 12; 107(9):3780-99; Available from: <https://doi.org/10.1021/cr068046j>.
- Quang DT, Kim JS. Fluoro- and Chromogenic Chemodosimeters for Heavy Metal Ion Detection in Solution and Biospecimens. Chemical Reviews. 2010 Oct 13; 110(10):6280-301; Available from: <https://doi.org/10.1021/cr100154p>.
- Hien NK, Bao NC, Ai Nhung NT, Trung NT, Nam PC, Duong T, et al. A highly sensitive fluorescent chemosensor for simultaneous determination of Ag(I), Hg(II), and Cu(II) ions: Design, synthesis, characterization and application. Dyes and Pigments. 2015 May 1; 116:89-96; Available from: <https://www.sciencedirect.com/science/article/pii/S0143720815000170>.
- Quy PT, Quang DT, Tung TQ. Design, synthesis and interpretation of fluorescent hemodosimeter based on density functional theory. Hue University Journal of Science: Natural Science. 2017 Apr 14; 126(1A):19-29; Available from: <https://doi.org/10.26459/hueuni-jns.v126i1A.3808>.
- Lutfullah, Sharma S, Rahman N, Azmi SNH, Iqbal B, Amburk MIBB, et al. UV Spectrophotometric Determination of Cu(II) in Synthetic Mixture and Water Samples. Journal of the Chinese Chemical Society. 2010 Aug; 57(4A):622-31; Available from: <https://doi.org/10.1002/jccs.201000087>.
- Souza JC TA, Beluomini MA, Eiras SP. Spectrophotometric determination of copper (II) in sugarcane spirit using 1-(2-pyridylazo)-2-naphthol and a homogeneous ternary mixture of the solvents water, ethanol and methyl isobutyl ketone. Re-

- 410 Vista Virtual De Quimica. 2016 Jan 1:687-701;Available from:
411 <http://hdl.handle.net/11449/163352>.
- 412 13. Omar AF, Atan H, MatJafri MZ. Peak Response Identification
413 through Near-Infrared Spectroscopy Analysis on Aqueous Su-
414 crose, Glucose, and Fructose Solution. Spectroscopy Letters.
415 2012 Apr 1; 45(3):190-201;Available from: [https://doi.org/10.](https://doi.org/10.1080/00387010.2011.604065)
416 [1080/00387010.2011.604065](https://doi.org/10.1080/00387010.2011.604065).
- 417 14. Tan CH, Moo YC, Jafri MZM, Lim HS. UV spectroscopy deter-
418 mination of aqueous lead and copper ions in water. ProcSPIE.
419 2014 May 15; 9141:91410 N;Available from: [https://doi.org/10.](https://doi.org/10.1117/12.2052349)
420 [1117/12.2052349](https://doi.org/10.1117/12.2052349).